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## Soap-Making Manual (Large Print) A Practical Handbook on the Raw Materials, Their Manipulation, Analysis and Control in the Modern Soap Plant

*E. G. Thomssen*

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**E. G. Thomssen : Soap-Making Manual (Large Print) A Practical Handbook on the Raw Materials, Their Manipulation, Analysis and Control in the Modern Soap Plant** before purchasing it in order to gauge whether or not it would be worth my time, and all praised Soap-Making Manual (Large Print) A Practical Handbook on the Raw Materials, Their Manipulation, Analysis and Control in the Modern Soap Plant:

New Edition, Large Print, 15 point font Soap is ordinarily thought of as the common cleansing agent well known to everyone. In a general and strictly chemical sense this term is applied to the salts of the non-volatile fatty acids. These

salts are not only those formed by the alkali metals, sodium and potassium, but also those formed by the heavy metals and alkaline earths. Thus we have the insoluble soaps of lime and magnesia formed when we attempt to wash in "hard water"; again aluminum soaps are used extensively in polishing materials and to thicken lubricating oils; ammonia or "benzine" soaps are employed among the dry cleaners. Commonly, however, when we speak of soap we limit it to the sodium or potassium salt of a higher fatty acid. It is very generally known that soap is made by combining a fat or oil with a water solution of sodium hydroxide (caustic soda lye), or potassium hydroxide (caustic potash). Sodium soaps are always harder than potassium soaps, provided the same fat or oil is used in both cases. The detergent properties of soap are due to the fact that it acts as an alkali regulator, that is, when water comes into contact with soap, it undergoes what is called hydrolytic dissociation. This means that it is broken down by water into other substances. Just what these substances are is subject to controversy, though it is presumed caustic alkali and the acid alkali salt of the fatty acids are formed.